

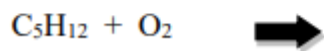
MINISTRY OF EDUCATION  
SECONDARY ENGAGEMENT PROGRAMME

GRADE 11  
CHEMISTRY

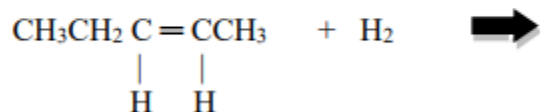
WEEK 5

LESSON 1

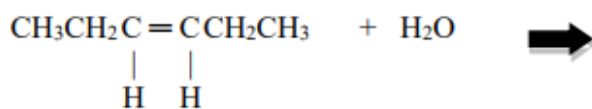
1 **combustion reaction** (assume "complete combustion")



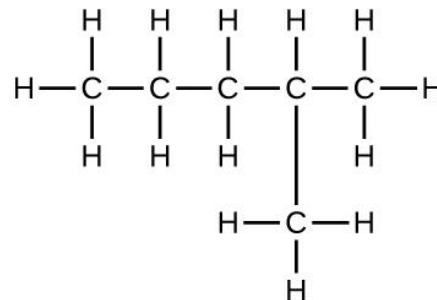
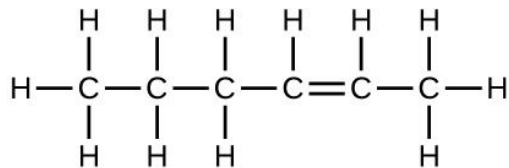
2.



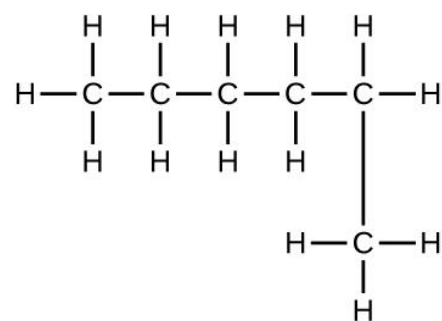
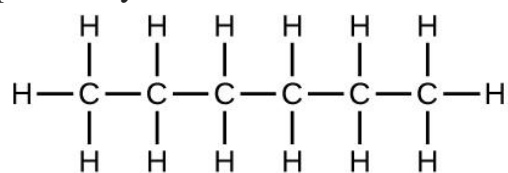
3.



4. Explain why these two molecules are not isomers:



5. Explain why these two molecules are not isomers:

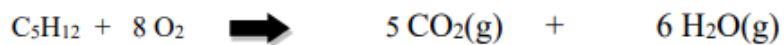


6.

Structural	Condensed
$  \begin{array}{ccc}  & \text{H} & \text{H} & \text{H} \\  &   &   &   \\  \text{H} & -\text{C} & -\text{C} & -\text{C}-\text{H} \\  &   &   &   \\  & \text{H} & \text{H} & \text{H}  \end{array}  $ <p style="text-align: center;">Propane</p>	
	<p style="text-align: center;"><math>\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3</math></p> <p style="text-align: center;">Hexane</p>
$  \begin{array}{cccc}  & \text{H} & \text{H} & \text{H} & \text{H} \\  &   &   &   &   \\  \text{H} & -\text{C} & =\text{C} & -\text{C} & -\text{C}-\text{H} \\  &   & &   &   \\  & \text{H} & & \text{H} & \text{H}  \end{array}  $ <p style="text-align: center;">Butene</p>	
	<p style="text-align: center;"><math>\text{CH}_2\text{CHCH}_3</math></p> <p style="text-align: center;">Propene</p>

1.

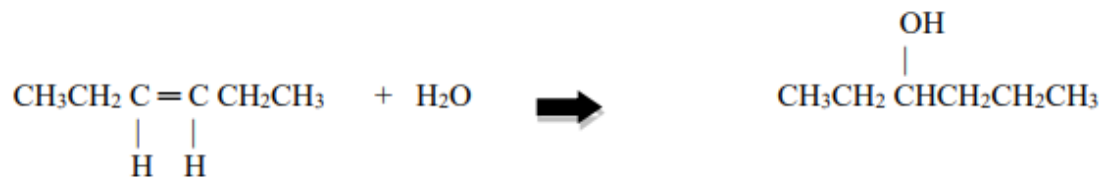
**combustion reaction** (assume "complete combustion"). The products for complete combustion are always  $\text{CO}_2(\text{g})$  and  $\text{H}_2\text{O}(\text{g})$ .



2.



3.



4. They are the same compound because each is a saturated hydrocarbon containing an unbranched chain of six carbon atoms.