

MINISTRY OF EDUCATION
SECONDARY ENGAGEMENT PROGRAMME
GRADE 10
CHEMISTRY

WEEK 10

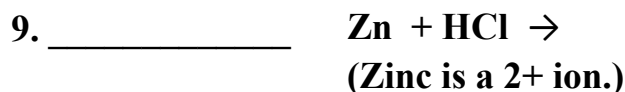
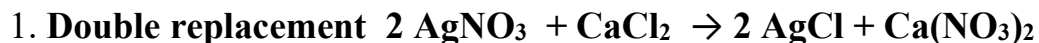
LESSON 1 - WORKSHEET

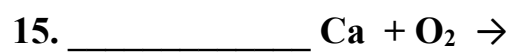
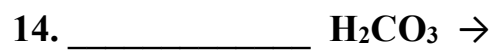
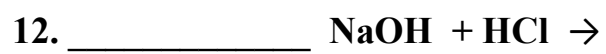
Predicting Products of Chemical reactions

Directions:

1. Determine the type of reaction
2. Predict the products in the reaction and write the correct chemical formula for each product
3. Balance each reaction using the correct coefficients.

Example





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LESSON 1 – ANSWERS

- 1. Double replacement $2 \text{AgNO}_3 + \text{CaCl}_2 \rightarrow 2 \text{AgCl} + \text{Ca}(\text{NO}_3)_2$**
- 2. Combination _____ $2 \text{Al} + 3 \text{Cl}_2 \rightarrow 2 \text{AlCl}_3$**
- 3. Double replacement $\text{FeCl}_2 + 2 \text{KOH} \rightarrow \text{Fe}(\text{OH})_2 + 2 \text{KCl}$**
- 4. Combustion _____ $\text{C}_2\text{H}_4 + 3 \text{O}_2 \rightarrow 2 \text{CO}_2 + 2 \text{H}_2\text{O}$**
- 5. Double replacement $\text{Al}_2(\text{SO}_4)_3 + 3 \text{Pb}(\text{NO}_3)_2 \rightarrow 2 \text{Al}(\text{NO}_3)_3 + 3 \text{PbSO}_4$**
- 6. Single replacement $\text{Ba} + \text{MgI}_2 \rightarrow \text{BaI}_2 + \text{Mg}$**
- 7. Decomposition ___ $\text{Ca}(\text{OH})_2 + \text{heat} \rightarrow \text{CaO} + \text{H}_2\text{O}$**
- 8. Double replacement $\text{H}_2\text{SO}_4 + 2 \text{KOH} \rightarrow \text{K}_2\text{SO}_4 + 2 \text{H}_2\text{O}$**
- 9. Single replacement $\text{Zn} + 2 \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$**
- 10. Combination/Synthesis _____ $2 \text{K} + \text{Br}_2 \rightarrow 2 \text{KBr}$**
- 11. Single replacement $\text{Mg} + \text{FeSO}_4 \rightarrow \text{MgSO}_4 + \text{Fe}$**
- 12. Double replacement $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$**
- 13. Combustion _____ $\text{C}_5\text{H}_{12} + 8 \text{O}_2 \rightarrow 5 \text{CO}_2 + 6 \text{H}_2\text{O}$**
- 14. Decomposition ___ $\text{H}_2\text{CO}_3 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$**
- 15. Combination _____ $2 \text{Ca} + \text{O}_2 \rightarrow 2 \text{CaO}$**
- 16. Single replacement $2 \text{Li} + \text{Cu}(\text{NO}_3)_2 \rightarrow \text{Cu} + 2 \text{LiNO}_3$**